

Figure 1. The effect of the concentration of the *Ag* ions on the  $\log K_{sp}$  of the  $\text{Ag}_2\text{S}$  and  $\text{Ag}_2\text{Se}$  precipitates. The  $\log K_{sp}$  values were calculated from the  $\log K_{sp}$  values of the  $\text{Ag}_2\text{S}$  and  $\text{Ag}_2\text{Se}$  precipitates at the concentration of the  $\text{Ag}^+$  ions of 0.01 M. The  $\log K_{sp}$  values of the  $\text{Ag}_2\text{S}$  and  $\text{Ag}_2\text{Se}$  precipitates at the concentration of the  $\text{Ag}^+$  ions of 0.01 M were 1.5 and 1.0, respectively. The  $\log K_{sp}$  values of the  $\text{Ag}_2\text{S}$  and  $\text{Ag}_2\text{Se}$  precipitates at the concentration of the  $\text{Ag}^+$  ions of 0.01 M were 1.5 and 1.0, respectively. The  $\log K_{sp}$  values of the  $\text{Ag}_2\text{S}$  and  $\text{Ag}_2\text{Se}$  precipitates at the concentration of the  $\text{Ag}^+$  ions of 0.01 M were 1.5 and 1.0, respectively.

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